Chemistry mole conversions worksheet

Continue

Mole Conversions Practice **NAME** _____

There are three mole equalities. Use your road map. They are:
1 mol = 6.02 x 10^{23} particles [atoms, molecules, formula units]
1 mol = Molar Mass [also called gram-formula-mass]
1 mol = 22.4 L for a gas at STP

Mole-Particle Conversions SHOW ALL WORK
EX. How many moles of magnesium is 3.01 x 10^{22} atoms of magnesium?

3.01 x 10^{22} atoms $\left(\frac{1\ mole}{6.02\times10^{23}\ atoms}\right)$ = 5 x 10^{-2} moles of magnesium

1. How many molecules are there in 4.00 moles of glucose, $C_6H_{12}O_6$?

2. How many moles are 1.20 x 10^{25} atoms of phosphorous?

3. How many atoms are in 0.750 moles of zinc?

4. How many molecules are in 0.400 moles of N_2O_5?

Mole-Mass Conversions SHOW ALL WORK
Ex. How many moles in 28 grams of CO_2 ?

Molar Mass of CO_2 C : 1 x 12.01 g = 12.01 g
O : 2 x 16.00 g = 32.00 g +
44.01 g/mol

28 g CO_2 $\left(\frac{1\ mole}{44.01\ g}\right)$ = 0.64 moles CO_2

1. What is the mass of 5 moles of Fe_2O_3 ?

Chemistry Worksheet: Mole Concept

Name _____ Date _____ Class Period _____

1) What is the mass of one mole of Br atoms?

2) What is the mass of one mole of titanium atoms?

3) What is the molar mass of tin?

4) What is the mass of one mole of chromium atoms?

5) What is the mass of 6.02 x 10^{23} atoms of arsenic?

6) Is one move of zinc the same as one atom of zinc?

7) How many atoms are there in one mole of atoms?

8) What is the mass of one mole of Br_2 atoms?

9) What is the mass of one mole of water?

10) What is the molar mass of $CaCl_2$?

11) What is the molar mass of $K_2Cr_2O_7$?

12) What is the mass of one mole of $Ba(OH)_2$?

13) What is the molar mass of carbon disulfide?

14) What is the molar mass of ethanol, C_2H_5OH?

15) Calculate the molar mass of each of the following:
a) $(NH_4)_2SO_3$ b) $Ca_3(PO_4)_2$ c) $Al_2(CrO_4)_3$ d) KNO_3

16) What is the mass of 6.02 x 10^{23} molecules of H_2CO_3?

mole $\left(\frac{gram}{1\ mole}\right)$ = gram gram $\left(\frac{1\ mole}{gram}\right)$ = mole

Mole Calculation Worksheet

1) How many moles are in 15 grams of lithium?
15g Li $\left(\frac{1\ mole}{7\ g}\right)$ = 2.14 MOL Li

2) How many grams are in 2.4 moles of sulfur?
2.4 MOL S $\left(\frac{32\ g\ S}{1\ mol\ S}\right)$ = 76.8 g S

3) How many moles are in 22 grams of argon?
22 g Ar $\left(\frac{1\ mole}{40\ g\ Ar}\right)$ = 0.55 MOL Ar

4) How many grams are in 88.1 moles of magnesium?
88.1 mol Mg $\left(\frac{24\ g\ Mg}{1\ MOL\ Mg}\right)$ = 2114.4 g Mg

5) How many moles are in 2.3 grams of phosphorus?
2.3 g P $\left(\frac{1\ mole\ P}{31\ g\ P}\right)$ = 0.07 MOL P

6) How many grams are in 11.9 moles of chromium?
11.9 MOL Cr $\left(\frac{52\ g\ Cr}{1\ mol\ Cr}\right)$ = 618.8 g Cr

7) How many moles are in 9.8 grams of calcium?
9.8 g Ca $\left(\frac{1\ mole\ Ca}{40\ g\ Ca}\right)$ = 0.245 mol Ca

8) How many grams are in 238 moles of arsenic?
238 MOL As $\left(\frac{75\ g\ As}{1\ mol\ As}\right)$ = 17,850 g As

What are the molecular weights of the following compounds?

9) NaOH
23 + 16 + 1
40 g

10) H_2O
1(2) + 16
18 g

11) $MgCl_2$
24 + 35(2)
94 g

12) H_3PO_4
1(3) + 31 + 16(4)
98 g

13) Mn_2Se_7
55(2) + 79(7)
663 g

14) $(NH_4)_2SO_4$
14(2) + 1(8) + 32 + 16(4)
28 + 8 + 32 + 64
132 g

Answers Not Rounded to Sig Figs Key

AMU Worksheet -

1. How many amu in 4.2 g of sulfur?
$\frac{4.2\ g}{1.66\times10^{-24}\ g} \left|\frac{1\ amu}{}\right. = 2.530120482\times10^{24}\ amu$

2. How many grams in 1.20 x 10^{35} amu of carbon?
$\frac{1.20\times10^{35}\ amu}{1} \left|\frac{1.66\times10^{-24}\ g}{1\ amu}\right. = 1.992\times10^{11}\ g$

3. How many atoms in 3.91 amu of sodium?
$\frac{3.91\ amu}{22.99\ amu} \left|\frac{1\ atom}{}\right. = 0.170073945\ atoms$

4. How many atoms in 45.22 amu of strontium?
$\frac{45.22\ amu}{87.62\ amu} \left|\frac{1\ atom}{}\right. = 0.516092216\ atoms$

5. How many molecules in 3.90 x 10^{15} amu of chlorine (Cl_2) gas? 35.45 x 2 = 70.90
$\frac{3.90\times10^{15}\ amu}{70.90\ amu} \left|\frac{1\ molecule}{}\right. = 5.500705219\times10^{13}\ molecules$

6. How many nitrogen atoms in 140 amu of nitrogen?
$\frac{140\ amu}{14.007\ amu} \left|\frac{1\ atom}{}\right. = 9.995002499\ atoms$

7. What mass would 4.52 x 10^{24} molecules of water have in AMU? H 1.008 x 2 = 2.016
O 15.999 x 1 = 15.999
$\frac{4.52\times10^{24}\ molecules}{1\ molecule} \left|\frac{18.015\ amu}{}\right. = 8.14278\times10^{25}\ amu$ 18.015

8. How many amu in 2.3 x 10^{-8} g of oxygen?
$\frac{2.3\times10^{-8}\ g}{1.66\times10^{-24}\ g} \left|\frac{1\ amu}{}\right. = 1.385542169\times10^{16}\ amu$

9. What is the mass of 1.2 x 10^{-3} moles of sodium chloride in amu? Na 22.99 x 1 = 22.99
Cl 35.45 x 1 = 35.45
55.44
$\frac{1.2\times10^{-3}\ moles}{1\ mol} \left|\frac{6.022\times10^{23}\ molecules}{}\right. = 7.2264\times10^{20}\ molecules \left|\frac{58.44\ amu}{1\ molecule}\right. = 4.223108816\times10^{22}\ amu$

10. How many amu of NaCl are there in 3.55 x 10^{-6} g of salt?
$\frac{3.55\times10^{-6}\ g}{1.66\times10^{-24}\ g} \left|\frac{1\ amu}{}\right. = 2.138554217\times10^{18}\ amu$

Chemistry 110
Practice Problems: gram/mole conversions

Avogadro's Number: $N_A = 6.02 \times 10^{23}\ /mole$

1. In 3 moles of O_2:
(a) How many O_2 molecules are there?

$$3 moles \times \frac{6.02 \times 10^{23}\ O_2\ molecules}{mole} = 1.8 \times 10^{24}\ O_2\ molecules$$

(b) How many O atoms are there?

$$3 moles \times \frac{6.02 \times 10^{23}\ O_2\ molecules}{mole} \times 2\ O\ atoms/O_2\ molecule = 3.6 \times 10^{24}\ O\ atoms$$

2. Calculate the number of moles in 40.0 g of carbon (C).

$$40.0 g \times \frac{mole}{12.01\ g} = 3.33\ moles$$

3. How many carbon atoms are there in 40.0 g?

$$3.33\ moles \times \frac{6.02 \times 10^{23}\ C\ atoms}{mole} = 2 \times 10^{24}$$

4. Calculate the number of moles in 40.0 g of carbon monoxide (CO).

$$molar\ mass = 12.01\ g/mol + 16.00\ g/mol = 28.01\ g/mol$$

$$40.0 g \times \frac{mole}{28.01\ g} = 1.43\ moles$$

5. How many grams are in 1.25 moles of neon (Ne)?

$$1.25\ moles \times \frac{20.0\ g}{mole} = 25\ g$$

6. How many grams are in 1.25 moles of gold (Au)?

$$1.25\ moles \times \frac{mole}{196.97\ g} = 246\ g$$

7. What is the molar mass of carbon dioxide (CO_2)?

$$molar\ mass = 12.01\ g/mol + 2(16.00\ g/mol) = 44.01\ g/mol$$

8. How many moles of CO_2 are there in 100.0 g?

$$100.0\ g \times \frac{mole}{44.01\ g} = 2.27\ moles$$

9. Calculate the molar mass of N_2H_4.

$$molar\ mass = 2(14.01\ g/mol) + 4(1.01\ g/mol) = 32.06\ g/mol$$

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Color Printable Periodic Table - Color table with atomic numbers, element symbols, atomic weights, periods, and groups. (no names) Basic Printable Periodic Table - Black/white table with atomic numbers, element symbols, atomic weights, periods. (no names) Basic Periodic Table with Element Names - Black/white table with element symbols, names, atomic numbers, and periods. (no weights) Basic Periodic Table with Element Names (color) - Color periodic table with element symbols, names, atomic numbers, periods, and groups. (no weights) The atomic weights given on these tables are the most recent (2007) values as accepted by the IUPAC. This is a flow chart of the steps of the scientific method, available as a PDF file: Also available is a PDF of a pie chart of the elemental composition of the human body. In order to continue enjoying our site, we ask that you confirm your identity as a human. Thank you very much for your cooperation. Loading... Normal Solution Concentration Calculator In some cases, particularly in situations involving acid-base chemistry, the solution concentration is expressed in normality (N or CN). Normality is defined as the number of equivalent weights (or simply equivalents, eq) of solute dissolved per liter of solution (equivalents/L = N) (Equation 1). Normality is used in place of molarity because often 1 mole of acid does not neutralize 1 mole of base. Thus, in order to have a one-to-one relationship between acids and bases, many chemists prefer to express the concentration of acids and bases in normality. Similar to molarity, where the molecular weight (MW) is used for calculating the molar concentration, for normal solution concentrations, it is desirable to use the equivalent weight (EW). For acids, the equivalent weight (EW) is defined as the mass (in grams) of an acid that provides exactly 1 mole of hydrogen ions (H+) (to be precise, hydronium ions, H3O+) to a reaction. Another way of expressing this is that the equivalent weight is defined as the mass (in grams) of an acid that reacts with exactly 1 mole of hydroxide ions (OH−). For bases, the equivalent weight is defined as the mass (in grams) of a base that provides exactly 1 mole of hydroxide ions (OH−) to a reaction. Another way of expressing this is that the equivalent weight is defined as the mass (in grams) of a base that reacts with exactly 1 mole of hydrogen ions (H+). Because 1 mole of acid or base is equal to 1 equivalent of acid or base, equivalent weight is expressed in grams per equivalent (g/equivalent or g/eq) (Equation 2). While it is more appropriate to refer to equivalent weight as equivalent mass, this term is rarely used. (Equation 1) As stated above, the equivalent weight (EW, expressed in grams per equivalent or g/eq) is related to the chemical nature of the acid or base under consideration. It is calculated by dividing the molecular weight (MW) of an acid or base by the number of equivalents per mole for that acid or base (Equation 2). For acids, the number of equivalents per mole is the number of moles of hydrogen ions (H+) contributed to the reaction per mole of acid. For bases, the number of equivalents per mole is the number of hydroxide ions (OH−) contributed to the reaction per mole of base. Thus, knowledge of the chemical nature of the acid and/or base involved in the chemical reaction is necessary to determine the equivalent weight of the acid or base. Tables 1 and 2 provide a few examples of acids (Table 1) and bases (Table 2), the number of equivalents per mole for each acid or base, and the corresponding equivalent weight values. (Equation 2) Table 1. A few examples of acids and their values for molecular weight, number of equivalents per mole of acid, and equivalent weight Note: For an acid, the number of equivalents per mole of acid is the number of moles of hydrogen ion (H+) provided by 1 mole of acid in a reaction, or the number of moles of hydroxide ions (OH−) neutralized by 1 mole of acid in a reaction. Table 2. A few examples of bases and their values for molecular weight, number of equivalents per mole of base, and equivalent weight Note: For a base, the number of equivalents per mole of base is the number of moles of hydroxide ions (OH−) provided by 1 mole of base in a reaction, or the number of moles of hydrogen ions (H+) neutralized by 1 mole of base in a reaction. Given the above information, the normal concentration of a solution can be calculated by using Equation 3, where CN is the normality or normal concentration, m is the mass of acid or base used, V is the total solution volume, and EW is the equivalent weight. (Equation 3) The normal concentration of a solution (normality, CN) is always equal to or greater than the molar concentration (molarity, CM) of the solution. The normal concentration can be calculated by multiplying the molar concentration by the number of equivalents per mole of solute (Equation 4). See Tables 1 and 2 for some typical values for the number of equivalents per mole of acid (Table 1) or base (Table 2). (Equation 4) It should also be noted that normality and equivalents are not used only in acid-base chemistry, but also in other applications such as reduction-oxidation (redox) reactions. Finally, it should be noted that the concept of normality evolved before the concept of molarity. Today, normality is not as widely used as it was by chemists a generation ago. Meant to be used in both the teaching and research laboratory, this calculator (see below) can be utilized to perform a number of different calculations for preparing normal solutions when starting with the solid material. For example, the equivalent weight of a chemical can be

used along with the desired solution volume and solute concentration to determine the mass of chemical needed to make such a solution. On the other hand, if the desired concentration is known, but only a small amount (i.e., mass) of the chemical is available (such as when a very small amount is purchased, e.g., 10 mg), then the total solution volume can be calculated in which the solid material must be dissolved to achieve the desired final concentration. Normal solution concentration equation CN is the normal concentration in equivalents/L (Normal or N). This is also referred to as normality, which is used to express the concentration of a solute in a solution. Normality is defined as the number of equivalents of solute dissolved per liter of solution (equivalents/L = N) (Equations 1, 3, and 4). A 1 N solution is one in which exactly 1 equivalent of solute is dissolved in a total solution volume of exactly 1 L. The normal concentration (CN) is always equal to or greater than the molar concentration (CM) (Equation 3). Using SI prefixes, the concentration may also be expressed in different fractions of the normal concentration such as milliequivalents/L (mN), microequivalents/L (µN), nanoequivalents/L (nN), picoequivalents/L (pN), etc. m is the mass (i.e., weight) of solute in grams (g) that must be dissolved in volume V of solution to make the desired normal concentration (CN). V is volume of solution in liters (L) in which the indicated mass (m) of solute must be dissolved to make the desired normal concentration (CN). Note that V is the final or total volume of solution after the solute has been added to the solvent. EW is the equivalent weight in g/equivalent. It is calculated by dividing the molecular weight of solute by the number of equivalents per mole of solute (Equation 2). For acids, the number of equivalents per mole is the number of H+ ions contributed by the acid per mole of acid. For bases, the number of equivalents per mole is the number of OH− ions contributed by the base per mole of base. The molecular weight of an acid or base can be obtained from the molecular formula, data tables, or the label on the bottle containing the acid or base of interest. Normal solution concentration calculator Each calculator cell shown below corresponds to a term in the formula presented above. Enter appropriate values in all cells except the one you wish to calculate. Therefore, at least three cells must have values, and no more than one cell may be blank. The value of the blank cell will be calculated based on the other values entered. After a calculation is performed, the calculated cell will be highlighted and subsequent calculations will calculate the value of the highlighted cell (with no requirement to have a blank cell). However, a blank cell has priority over a highlighted cell. For convenience, this worksheet allows you to select different mass, volume, and concentration units, and the necessary conversions are carried out for you to obtain the value of the blank cell in the desired unit. Note that the unit of equivalent weight must be g/equivalent. Comments and/or preparation instructions Posted: Sunday, April 10, 2016Last updated: Saturday, December 23, 2017 You can take your vehicle to a CarMax location to have it inspected or get a quote online if your vehicle is eligible. If it passes muster, you'll get an instant cash offer. That offer is good for seven days from the time you get it. If you do sell your car to CarMax, you're under no obligation to buy a. Does Carmax Register My Car will sometimes glitch and take you a long time to try different solutions. LoginAsk is here to help you access Does Carmax Register My Car quickly and handle each specific case you encounter. 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We've helped clients across the country get fast and easy quotes for their damaged vehicles and get paid fast. At CarBrain, we know you're too busy to deal with scammers, swindlers and expensive car problems. Our promise to you is transparency, convenience and. The vehicle had been cleaned and prepped and I could see the oil was new and clean and remains so given the miles driven so far. The OCI and service intervals have all been reset to 10K. Car is a 2016 Lexus Es 350 23K at time of purchase. Do they use bulk Dino, synthetic blend? Full syn?. Step 3: Apply for Registration at the DMV. After these steps, you're ready to register your car. Head to the DMV and submit your most recent registration, out-of-state license plates, and vehicle title (unless held by a lienholder). With these, the Application for Vehicle Registration, and some payments (the registration fee** and government. CarMax will buy most cars, even old cars that dont even run. However, these cars will need to be towed to a CarMax store for an appraisal. CarMax will even buy. Your Carmax representative may offer you wholesale price in an effort to make maximum profit on your vehicle. Chances are, selling it is the only reasonable option. There are many places that will buy junk cars, but something many people find themselves wondering is does CarMax buy cars that don't run. Simply put, yes, they do. However, there is some required documentation that's needed before you can do that. This includes:. A DMV vehicle history report helps you save money, get the best deal and make a smart and informed choice. If the car was used as rental or a taxicab, it was most likely abused. Similarly a car which was previously in use of the police department would have been situations which may have made accidents more likely. A DMV vehicle history report. The possibilities are endless. 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You can sell a car fast, get a free quote, or call us (224) 236-7010. We will then reach out to you with an offer. If you agree with it, we will schedule you for a visit to inspect your car. After inspection, we will let you sign some paperwork and we will pay you with the agreed amount. However, when the car gets to the requested location, CarMax won't force you to buy it. 2. State fees. Examples of state fees are registration fees, sales tax, title transfers, and others. They vary from state to state. You can visit the nearest CarMax or your state's DMV website to learn about the various state fees you need to pay. If your car doesn't pass smog you will be unable to register the car. This means that you are not legally able to drive it on the city streets. This is handled by the department of motor vehicles, and there are no exceptions to the rule. The goal here is to keep the environment clean, the state is only doing this to preserve the environment. 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